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by diaphragm

cell) ~~Chemical~~

~~reaction \u0026~~

~~Equation~~


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\u0026 Kinetic

Friction,

Tension, Normal

Force, Inclined

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concentration of

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of NaOH by

diaphragm cell)

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Chapter 3 Quiz.

True/False.

Indicate whether

the sentence or

statement is

true or false.

1. The region in

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space where an electron is most likely to be found is called an energy level.

2. Electron configurations are often condensed by writing them using the previous noble gas core as a starting point.

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Solution: OH OH
OH 0.50 mmol mL
-- --x = nc V
x25.0 mL OH n -~~

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=12.5 mmol

HCl(aq) HCl(aq)

[HCl(aq)] 12.5

mmol = = n V

0.10 mmol

HCl(aq) /mL V

=120 mL 2. (a)

Given: V HCl(aq)

= 25.00 mL;

[HCl(aq)] =

0.350 mol/L; V

NaOH(aq) = 5.0

mL; [NaOH(aq)] =

0.500 mol/L;

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Required:

$[H^+(aq)]$, pH

Analysis: Before

titration,

amount of H^+ is:

$H(aq)$ $H(aq)$

$H(aq)$ 0.350 mmol

mL $++$ $+ = x = nc V$

~~Section 8.7:~~

~~Acid-Base~~

~~Titration~~

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Identify any errors in the structure by drawing them.

Rename the structure correctly. 2-ethyl-2,4,4-trimethylpentane What Is Required? You are to identify

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the errors in the name by drawing the structure and renaming it correctly. What Is Given? You are given an incorrectly named alkane.

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$$= 0.12 \text{ mol/L}$$

Statement: The equilibrium concentration of dinitrogen tetroxide gas is 0.12 mol/L. 2.

Given: initial quantity of $\text{NOCl(g)} = 3.00$ mol; volume = 3.00 L; $[\text{NO(g)}]$ initial = 0.00

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mol/L; $[\text{Cl}_2(\text{g})]$

initial = 0.00

mol/L; $[\text{NO}(\text{g})]$

equilibrium =

0.043 mol/L

Required:

$[\text{NOCl}(\text{g})]$

equilibrium; $[\text{Cl}_2(\text{g})]$

$[\text{Cl}_2(\text{g})]$

equilibrium

Solution: Step

1.

~~Section 7.1:~~

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~~Equilibrium
Systems Mini
Investigation:
The ...~~

Solution: bonds
broken 3 mol 1
mol CH CCl 3 mol
HnD DD -- $\Delta = x$
 $= x + x = \sum 413 \text{ kJ}$
mol \times +1 mol 339
kJ mol \times ΔH
 $= 1578 \text{ kJ}$

Statement: The
energy required

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to separate 1 mol of chloromethane into free atoms is 1578 kJ. 2. Solution: Step 1: Write the balanced chemical equation for the combustion of 1 mol of ethene gas. $C_2H_4(g) + 3 O_2(g) \dots$

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~~Section 5.3:
Bond Energies
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Solution: $q = Ne =$
 $(1.2 \times 10^8) (1.602 \times$
 $10^{-19} \text{ C})$

$q = 1.92 \times 10^{-11} \text{ C}$

(one extra digit
carried) $F E =$

$kq_1 q_2 / r^2 =$

$8.99 \times 10^9 \text{ N} \cdot \text{m}^2 / \text{C}^2$

\$ % & ' "

$((1.92 \times 10^{-11})^2 /$

$(1.10 \text{ m})^2) = 1.4 \times 10^{-21} \text{ N}$

$(1.10 \text{ m})^2 = 1.21 \text{ m}^2$

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$= 2.7 \times 10^{-12}$ N

Statement: The force of repulsion

between the two plastic spheres is 2.7×10^{-12}

N. 2. Given: $m = 2.48 \times 10^{-15}$ kg;

$\Delta d = 1.7$ cm =

0.017 m; ΔV b =

260 V; $e = 1.602$

$\times 10^{-19}$ C; $g =$

9.8 m/s²

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